Basic Chemistry 6th edition Timberlake Test Bank

SKU: 9780134987088-TEST-BANK

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

1) 5.21 cm is the same distance as _____.
A) 0.000 521 km
B) 5210 m
C) 0.0521 m
D) 5.21 mm
E) 52.1 dm

Answer: C

2) The measurement of the gravitational pull on an object is its _____.
A) weight
B) volume
C) mass
D) length
E) size

Answer: A

3) The amount of space occupied by a substance is its _____.
A) volume
B) density
C) mass
D) weight
E) length

Answer: A

4) Which of the following is the basic unit of volume in the metric system?
A) kilogram
B) liter
C) gram
D) centimeter
E) meter

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Answer: B

5) Which of the following is the SI unit of mass?
A) milliliter
B) kilogram
C) centimeter
D) Celsius
E) meter

Answer: B

6) A value of 25 °C is a measurement of _____.
A) volume
B) temperature
C) density
D) distance
E) mass

Answer: B

T B E X A M . C O M

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7) Which of the following conversion factors involves a measured number?
A) 12 in/ft
B) 10 cm/dm
C) 16 oz/lb
D) 12 eggs/dozen
E) 25 miles/gallon

Answer: E

8) Which of the following measured numbers has three significant figures?
A) 0.001 cm
B) 1.01 cm
C) 10.01 cm
D) 1.0×10^3 cm
E) 100 cm

Answer: B

9) Which of the following measured numbers has two significant figures?
A) 200 cm
B) 2.0×10^3 mL
C) 0.002 mL
D) 0.2 mL
E) 20.0 mL

Answer: B

10) Significant figures are important because they indicate _____.
A) the number of measurements
B) a counted number
C) the accuracy of the conversion factor
D) the number of digits on a calculator
E) the number of digits in a measurement

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Answer: E

11) Which of the following measurements has three significant figures?
A) 0.510 m
B) 0.005 m
C) 5100 m
D) 510 m
E) 0.051 m

Answer: A

12) Which of the following numbers contains the designated CORRECT number of significant figures?
A) 1.04 2 significant figures
B) 156 000 3 significant figures
C) 0.043 00 5 significant figures
D) 3.0650 4 significant figures
E) 0.00302 2 significant figures

Answer: B

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13) The number of significant figures in the measurement of 45.030 mm is _____.
A) none
B) five
C) three
D) six
E) four

Answer: B

14) How many significant figures are in the number 0.00208?
A) six
B) five
C) two
D) three
E) four

Answer: D

15) Which of the following examples illustrates a number that is correctly rounded to three significant figures?
A) 0.03954 grams to 0.040 grams
B) 109 526 grams to 109 500 grams
C) 4.05438 grams to 4.054 grams
D) 20.0332 grams to 20.0 grams
E) 103.692 grams to 103.7 grams

Answer: D

16) A calculator answer of 423.6059 must be rounded off to three significant figures. What answer is reported?
A) 423.7
B) 423
C) 423.6
D) 420
E) 424

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Answer: E

17) Which of the answers for the following conversions contains the correct number of significant figures?

A) $12.0 \text{ ft} \times \frac{12 \text{ in.}}{1 \text{ ft}} \times \frac{2.54 \text{ cm}}{1 \text{ in}} = 370 \text{ cm}$

B) $2.543 \text{ m} \times \frac{39.37 \text{ in}}{1 \text{ m}} = 100.1942 \text{ in}$

C) $2 \text{ L} \times \frac{1.057 \text{ qt}}{1 \text{ L}} = 2.12 \text{ qt}$

D) $24.0 \text{ kg} \times \frac{1 \text{ lb}}{2.205 \text{ kg}} = 11 \text{ lb}$

E) $24.95 \text{ min} \times \frac{1 \text{ h}}{60 \text{ min}} = 0.4158 \text{ h}$

Answer: E

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18) What is the correct answer for the calculation of a volume (in mL) with measured numbers $\frac{28.58}{16 \times 8.02}$?

A) 0.22 mL
B) 14.3 mL
C) 57 mL
D) 14 mL
E) 0.223 mL

Answer: A

19) A researcher needed three samples of sodium chloride solution, each with a volume of 0.03510 mL. The total volume needed, if the three volumes are added together, should be reported as _____.

A) 0.10 mL
B) 0.1053 mL
C) 0.0105 mL
D) 0.10530 mL
E) 0.105 mL

Answer: D

20) What is the answer, with the correct number of significant figures, for this problem?

4.392 g + 102.40 g + 2.51 g =

A) 109.30 g
B) 110 g
C) 109.302 g
D) 109.3 g
E) 109 g

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Answer: A

21) The correct answer for the addition of 7.5 g + 2.26 g + 1.311 g + 2 g is _____.

A) 13 g
B) 13.0 g
C) 10 g
D) 13.1 g
E) 13.071 g

Answer: A

22) In which of the following is the metric unit paired with its correct abbreviation?

A) milliliter /mL
B) kilogram /cg
C) microgram /mg
D) gram /gm
E) centimeter /km

Answer: A

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23) Which of the following measurements are NOT equivalent?
A) 183 L = 0.183 kL
B) 24 dL = 2.4 L
C) 150. msec = 0.150 sec
D) 25 mg = 0.025 g
E) 84 cm = 8.4 mm

Answer: E

24) Which of the following is the largest unit?
A) millimeter
B) kilometer
C) micrometer
D) meter
E) decimeter

Answer: B

25) What is the metric relationship between grams and micrograms?
A) 1 g = 100 µg
B) 1 g = 0.000 001 µg
C) 1 g = 1 000 000 µg
D) 1 g = 0.001 µg
E) 1 g = 1000 µg

Answer: C

26) What is the conversion factor for the relationship between millimeters and centimeters?
A) 100 mm/1 cm
B) 10 mm/1 cm
C) 10 cm/1 mm
D) 1 cm/1 mm
E) 1 mm/1 cm

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Answer: B

27) Which of the following is the smallest unit?
A) milligram
B) kilogram
C) gram
D) decigram
E) microgram

Answer: E

28) The cubic centimeter (cm^3 or cc) has the same volume as a _____.
A) centimeter
B) cubic inch
C) cubic decimeter
D) cubic liter
E) milliliter

Answer: E

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29) 9.31 g is the same mass as _____.
A) 931 kg
B) 93.1 cg
C) 0.0931 dg
D) 931 µg
E) 9310 mg

Answer: E

30) An alloy of iron contains 75.0% iron and 25.0% other elements. How many grams of iron are present in 150. g of the alloy?
A) 37.5 g
B) 11 300 g
C) 3750 g
D) 113 g
E) 2.00 g

Answer: D

31) One form of stainless steel contains 18.0% nickel. How much nickel is present in 200. g of this alloy?
A) 18.0 g
B) 36.0 g
C) 0.0122 g
D) 164 g
E) 11.1 g

Answer: B

32) A 100.0 g sample of eighteen karat gold is contains 75.0 g of gold and 25.0 g of other metals. What is the percent of gold in the sample?
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A) 50%
B) 125%
C) 100.0%
D) 75.0%
E) 25.0%

Answer: D

33) An sample of hamburger had a total mass of 200. g, of which 30.0 g was found to be fat. What is the percent of fat in this hamburger sample?
A) 15.0%
B) 30.0%
C) 6.67%
D) 13.3%
E) 6.00%

Answer: A

34) What is the correct conversion factor for milligrams and micrograms?
A) 10 mg/1mcg
B) 1000 mg/1 mcg
C) 10^6 mcg/1 mg
D) 1000 mcg/1 mg
E) 1 mg/100 mcg

Answer: D

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35) What is the correct conversion factor for kilometers and millimeters?
A) 1 km/1000 mm
B) 1000 mm/1 km
C) 10^6 mm/1 km
D) 100 mm/1 km
E) 10^6 km/1 mm

Answer: C

36) According to the United States Food and Drug Administration, the recommended daily requirement of protein is 44 g. This is _____ oz of protein.
A) 150 000
B) 1248.5
C) 0.0605
D) 1.6
E) 320 000

Answer: D

37) Which of the following setups would convert centimeters to feet?
A) $\text{cm} \times \frac{1 \text{ in.}}{2.54 \text{ cm}} \times \frac{1 \text{ ft}}{12 \text{ in.}}$
B) $\text{cm} \times \frac{2.54 \text{ in.}}{1 \text{ cm}} \times \frac{1 \text{ ft}}{12 \text{ in.}}$
C) $\text{cm} \times \frac{2.54 \text{ cm}}{1 \text{ in.}} \times \frac{1 \text{ ft}}{12 \text{ in.}}$
D) $\text{cm} \times \frac{2.54 \text{ cm}}{1 \text{ in.}} \times \frac{12 \text{ in.}}{1 \text{ ft}}$
E) $\text{cm} \times \frac{1 \text{ in.}}{2.54 \text{ cm}} \times \frac{12 \text{ in.}}{1 \text{ ft}}$

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Answer: A

38) A conversion factor set up correctly to convert 15 inches to centimeters is _____.
A) 1 cm/10 mm
B) 10 cm/1 inch
C) 100 cm/1 m
D) 2.54 cm/1 inch
E) 1 inch/2.54 cm

Answer: D

39) How many pounds are in 3.5 kg?
A) 1.59 lb
B) 7.70 lb
C) 7.7 lb
D) 0.629 lb
E) 1.6 lb

Answer: C

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40) How many liters of soft drink are there in 5.25 qt?
A) 5.0 L
B) 4.97 L
C) 55.7 L
D) 4950 L
E) 5.57 L

Answer: B

41) What is 6.5 m converted to inches?
A) 255.9 in
B) 1700 in
C) 39 in
D) 260 in
E) 1651 in

Answer: D

42) How many kilograms are in 30.4 lb?
A) 13.8 kg
B) 14 kg
C) 66.9 kg
D) 66.88 kg
E) 67 kg

Answer: A

43) A dose of aspirin of 5.0 mg per kilogram of body weight has been prescribed to reduce the fever of an infant weighing 8.5 pounds. The number of milligrams of aspirin that should be administered is _____.
A) 5.0 mg
B) 1.6 mg
C) 53 mg
D) 19 mg
E) 0.59 mg

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Answer: D

44) If 5.00 lb of potatoes costs $3.60, how much would 1.30 kilograms of potatoes cost?
A) $0.86
B) $3.97
C) $2.06
D) $10.30
E) $0.43

Answer: C

45) How many centimeters are there in 57.0 in?
A) 0.0445 cm
B) 145 cm
C) 140 cm
D) 22.4 cm
E) 22 cm

Answer: B

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46) If a car travels 23 miles on 1.0 gal of gas, how many liters of gasoline are needed for a 135 mile trip?
A) 22 L
B) 5.9 gal
C) 32 L
D) 14 L
E) 25 L

Answer: A

47) The mercury level in cod was measured at 0.11 ppm. How many mg of mercury are present in a 150 g serving of cod?
A) 0.14 mg
B) 0.17 mg
C) 150 mg
D) 0.11 mg
E) 0.017 mg

Answer: E

48) The herbicide level in the soil in a corn field was measured at 3.0 ppb. How many µg of herbicide are present in 1.0 lb of soil?
A) 0.44 µg
B) 3.0 µg
C) 1.4 µg
D) 0.7 µg
E) 4.5 µg

Answer: C

49) A nugget of gold with a mass of 521 g is added to 50.0 mL of water. The water level rises to a volume of 77.0 mL. What is the density of the gold?
A) 10.4 g/mL
B) 19.3 g/mL
C) 6.77 g/mL
D) 1.00 g/mL
E) 0.0518 g/mL

Answer: B

50) A solution has a density of 1.22 g/mL. What volume of the solution has a mass of 48.2 g?
A) 0.00253 mL
B) 1.22 mL
C) 58.8 mL
D) 39.5 mL
E) 49.4 mL

Answer: D

51) Which one of the following substances will float in gasoline, which has a density (d) of 0.66 g/mL?
A) aluminum (d = 2.70 g/mL)
B) table salt (d = 2.16 g/mL)
C) balsa wood (d = 0.16 g/mL)
D) mercury (d = 13.6 g/mL)
E) sugar (d = 1.59 g/mL)

Answer: C

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52) What is the mass of 2.00 L of a solution with a density of 1.15 g/mL?
A) 0.58 kg
B) 2.30 kg
C) 1.15 kg
D) 0.015 kg
E) 0.023 kg
Answer: B

53) Mercury has a density of 13.6 g/mL. How many milliliters of mercury have a mass of 0.35 kg?
A) 0.026 mL
B) 0.0257 mL
C) 26 mL
D) 25.7 mL
E) 4760 mL
Answer: C

54) What is the density of a substance with a mass of 45.00 g and a volume of 26.4 mL?
A) 1.7 g/mL
B) 0.59 g/mL
C) 0.587 g/mL
D) 45.0 g/mL
E) 1.70 g/mL
Answer: E

55) What is the mass of 53 mL of ethyl alcohol, which has a density of 0.79 g/mL?
A) 41.9 g
B) 67 g
C) 42 g
D) 53 g
E) 67.1 g
Answer: C

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56) The density of a solution is 1.18 g/mL, and its volume is 25.0 mL. The mass of the sample is _____.
A) .0472 g
B) 25.0 g
C) 29.5 g
D) 1.18 g
E) 21.2 g
Answer: C

57) Diamond has a density of 3.52 g/mL. What is the volume in cubic centimeters of a diamond with a mass of 15.1 g?
A) 53 cm^3
B) 0.233 cm^3
C) 4.29 cm^3
D) 53.2 cm^3
E) 4.3 cm^3
Answer: C

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58) The ratio of the mass of a substance to its volume is its _____.
A) weight
B) conversion factor
C) specific gravity
D) density
E) buoyancy
Answer: D

59) A 50.0 mL liquid sample has a mass of 50.7 g. The density of the sample is _____.
A) 1.01
B) 0.986 g/L
C) 1.01 g/mL
D) 50.7
E) 0.986
Answer: C

60) A solution has a specific gravity of 1.13. What is the mass of 36.6 mL of the solution?
A) 32.4 g
B) 36.6 g
C) 1.00 g
D) 41.4 g
E) 1.13 g
Answer: D

MATCHING. Choose the item in column 2 that best matches each item in column 1.

Are the numbers in each of the following statements measured or exact?

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61) In the U.S. system there are 5280 feet in one mile.
Answer: B

A) measured
B) exact

62) A lab test showed a blood sugar level is 350 mg/dL.
Answer: A

63) There are 452 pages in a book.
Answer: B

64) The rabbit weighs 2.5 pounds.
Answer: A

65) There are 100 aspirin in a bottle.
Answer: B

66) You feel ill and your temperature is 100.1 °F.
Answer: A

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Match the type of measurement to the unit given below.

67) milliliter
Answer: D

68) mm
Answer: E

69) gram
Answer: B

70) 125 K
Answer: C

71) kilometer
Answer: E

A) density

B) mass

C) temperature

D) volume

E) distance

Select the correct numerical prefix to complete the equality.

72) 1 g = _____ kg
Answer: E

73) 1 m = _____ mm
Answer: B

74) 1 cm = _____ mm
Answer: A

75) 1 dL = _____ mL
Answer: D

76) 1 mL = _____ cc
Answer: C

A) 10

B) 1000

C) 1

D) 100

E) 0.001

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TRUE/FALSE. Write 'T' if the statement is true and 'F' if the statement is false.

77) A kilogram is a unit of volume.
Answer: True ✔ False

78) A liter is a unit of volume.
Answer: ✔ True False

79) The measurement 1.230 cm has 4 significant figures.
Answer: ✔ True False

80) The measurement 0.03550 has 4 significant figures.
Answer: ✔ True False

81) When the measurement 3.32 cm is multiplied by the measurement 0.02 cm, the answer will have three significant figures.
Answer: True ✔ False

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82) When the measurement 13.36 cm is added to the measurement 0.02 cm, the answer will be 13.38 cm.
Answer: True False

83) A microgram is larger than a gram.
Answer: True False

84) A 1- cup measuring cup holds about 240 mL.
Answer: True False

85) One conversion factor for cm and m is 100 m/1 cm.
Answer: True False

86) One conversion factor for mL and L is 1000 mL/1 L.
Answer: True False

87) 10.5 in is the same distance as 4.13 cm.
Answer: True False

88) A fish that weighs 15.5 lb has a mass of 7.03 kg.
Answer: True False

89) Water (density = 1.00 g/mL) will float on hexane (density = 0.95 mL).
Answer: True False

90) The mass of 10.0 mL of water is approximately 10.0 kg.
Answer: True False

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SHORT ANSWER. Write the word or phrase that best completes each statement or answers the question.

Round off each of the following to three significant figures.

91) 504.85
Answer: 505

92) 8.3158
Answer: 8.32

93) 25 225
Answer: 25 200

94) 58.5422
Answer: 58.5

95) 0.003 4088
Answer: 0.00341

State the number of significant figures in each of the following measurements.

96) 0.705 m
Answer: 3

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97) 680 000 km
Answer: 2

98) 28.050 km
Answer: 5

99) 0.0005 L
Answer: 1

100) 75.00 m
Answer: 4

101) 2.043×10^4 mm
Answer: 4

102) 6.1×10^{-5} mL
Answer: 2

103) 9.00×10^6 g
Answer: 3

104) The unit of volume in the SI system is the _____.
Answer: cubic meter

105) The unit of mass in the metric system is the _____.
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Answer: gram

106) Ten karat gold is 41.7% gold. How many grams of pure gold are there in a ring made of 70.0 g of ten karat gold?
Answer: 29.2 g

107) To calculate the density of a solid object, two measurements are needed, its _____ and _____.
Answer: mass, volume

108) Rubbing alcohol (isopropyl alcohol) has a density of 0.79 g/mL. How many mL of isopropyl alcohol contain 45 g of alcohol?
Answer: 57 mL

109) The density of gold is 19.3 g/mL. How many grams of gold are in a medal that has a volume of 15.0 mL?
Answer: 290. g of gold

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Answer Key
Testname: UNTITLED2

1) C
ID: TestBank 2.1-1
Diff: 0
Objective: 2.1

2) A
ID: TestBank 2.1-2
Diff: 0
Objective: 2.1

3) A
ID: TestBank 2.1-3
Diff: 0
Objective: 2.1

4) B
ID: TestBank 2.1-4
Diff: 0
Objective: 2.1

5) B
ID: TestBank 2.1-5
Diff: 0
Objective: 2.1

6) B
ID: TestBank 2.1-6
Diff: 0
Objective: 2.1

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7) E
ID: TestBank 2.1-7
Diff: 0
Objective: 2.2

8) B
ID: TestBank 2.1-8
Diff: 0
Objective: 2.2

9) B
ID: TestBank 2.1-9
Diff: 0
Objective: 2.2

10) E
ID: TestBank 2.1-10
Diff: 0
Objective: 2.2

11) A
ID: TestBank 2.1-11
Diff: 0
Objective: 2.2

12) B
ID: TestBank 2.1-12
Diff: 0
Objective: 2.2

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Answer Key
Testname: UNTITLED2

13) B
ID: TestBank 2.1-13
Diff: 0
Objective: 2.2

14) D
ID: TestBank 2.1-14
Diff: 0
Objective: 2.2

15) D
ID: TestBank 2.1-15
Diff: 0
Objective: 2.3

16) E
ID: TestBank 2.1-16
Diff: 0
Objective: 2.3

17) E
ID: TestBank 2.1-17
Diff: 0
Objective: 2.3

18) A
ID: TestBank 2.1-18
Diff: 0
Objective: 2.3

19) D
ID: TestBank 2.1-19
Diff: 0
Objective: 2.3

20) A
ID: TestBank 2.1-20
Diff: 0
Objective: 2.3

21) A
ID: TestBank 2.1-21
Diff: 0
Objective: 2.3

22) A
ID: TestBank 2.1-22
Diff: 0
Objective: 2.4

23) E
ID: TestBank 2.1-23
Diff: 0
Objective: 2.4

24) B
ID: TestBank 2.1-24
Diff: 0
Objective: 2.4

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25) C
ID: TestBank 2.1-25
Diff: 0
Objective: 2.4

26) B
ID: TestBank 2.1-26
Diff: 0
Objective: 2.4

27) E
ID: TestBank 2.1-27
Diff: 0
Objective: 2.4

28) E
ID: TestBank 2.1-28
Diff: 0
Objective: 2.4

29) E
ID: TestBank 2.1-29
Diff: 0
Objective: 2.4

30) D
ID: TestBank 2.1-30
Diff: 0
Objective: 2.5

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31) B
ID: TestBank 2.1-31
Diff: 0
Objective: 2.5

32) D
ID: TestBank 2.1-32
Diff: 0
Objective: 2.5

33) A
ID: TestBank 2.1-33
Diff: 0
Objective: 2.5

34) D
ID: TestBank 2.1-34
Diff: 0
Objective: 2.5

35) C
ID: TestBank 2.1-35
Diff: 0
Objective: 2.5

36) D
ID: TestBank 2.1-36
Diff: 0
Objective: 2.6

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Answer Key
Testname: UNTITLED2

37) A
ID: TestBank 2.1-37
Diff: 0
Objective: 2.6

38) D
ID: TestBank 2.1-38
Diff: 0
Objective: 2.6

39) C
ID: TestBank 2.1-39
Diff: 0
Objective: 2.6

40) B
ID: TestBank 2.1-40
Diff: 0
Objective: 2.6

41) D
ID: TestBank 2.1-41
Diff: 0
Objective: 2.6

42) A
ID: TestBank 2.1-42
Diff: 0
Objective: 2.6

43) D
ID: TestBank 2.1-43
Diff: 0
Objective: 2.6

44) C
ID: TestBank 2.1-44
Diff: 0
Objective: 2.6

45) B
ID: TestBank 2.1-45
Diff: 0
Objective: 2.6

46) A
ID: TestBank 2.1-46
Diff: 0
Objective: 2.6

47) E
ID: TestBank 2.1-47
Diff: 0
Objective: 2.6

48) C
ID: TestBank 2.1-48
Diff: 0
Objective: 2.6

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Answer Key
Testname: UNTITLED2

49) B
ID: TestBank 2.1-49
Diff: 0
Objective: 2.7

50) D
ID: TestBank 2.1-50
Diff: 0
Objective: 2.7

51) C
ID: TestBank 2.1-51
Diff: 0
Objective: 2.7

52) B
ID: TestBank 2.1-52
Diff: 0
Objective: 2.7

53) C
ID: TestBank 2.1-53
Diff: 0
Objective: 2.7

54) E
ID: TestBank 2.1-54
Diff: 0
Objective: 2.7

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55) C
ID: TestBank 2.1-55
Diff: 0
Objective: 2.7

56) C
ID: TestBank 2.1-56
Diff: 0
Objective: 2.7

57) C
ID: TestBank 2.1-57
Diff: 0
Objective: 2.7

58) D
ID: TestBank 2.1-58
Diff: 0
Objective: 2.7

59) C
ID: TestBank 2.1-59
Diff: 0
Objective: 2.7

60) D
ID: TestBank 2.1-60
Diff: 0
Objective: 2.7

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Answer Key
Testname: UNTITLED2

61) B
ID: TestBank 2.2-1
Diff: 0
Objective: 2.2

62) A
ID: TestBank 2.2-2
Diff: 0
Objective: 2.2

63) B
ID: TestBank 2.2-3
Diff: 0
Objective: 2.2

64) A
ID: TestBank 2.2-4
Diff: 0
Objective: 2.2

65) B
ID: TestBank 2.2-5
Diff: 0
Objective: 2.2

66) A
ID: TestBank 2.2-6
Diff: 0
Objective: 2.2

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67) D
ID: TestBank 2.2-7
Diff: 0
Objective: 2.1

68) E
ID: TestBank 2.2-8
Diff: 0
Objective: 2.1

69) B
ID: TestBank 2.2-9
Diff: 0
Objective: 2.1

70) C
ID: TestBank 2.2-10
Diff: 0
Objective: 2.1

71) E
ID: TestBank 2.2-11
Diff: 0
Objective: 2.1

72) E
ID: TestBank 2.2-12
Diff: 0
Objective: 2.4

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Answer Key
Testname: UNTITLED2

73) B
ID: TestBank 2.2-13
Diff: 0
Objective: 2.4

74) A
ID: TestBank 2.2-14
Diff: 0
Objective: 2.4

75) D
ID: TestBank 2.2-15
Diff: 0
Objective: 2.4

76) C
ID: TestBank 2.2-16
Diff: 0
Objective: 2.4

77) FALSE
ID: TestBank 2.3-1
Diff: 0
Objective: 2.1

78) TRUE
ID: TestBank 2.3-2
Diff: 0
Objective: 2.1

79) TRUE
ID: TestBank 2.3-3
Diff: 0
Objective: 2.2

80) TRUE
ID: TestBank 2.3-4
Diff: 0
Objective: 2.2

81) FALSE
ID: TestBank 2.3-5
Diff: 0
Objective: 2.3

82) TRUE
ID: TestBank 2.3-6
Diff: 0
Objective: 2.3

83) FALSE
ID: TestBank 2.3-7
Diff: 0
Objective: 2.4

84) TRUE
ID: TestBank 2.3-8
Diff: 0
Objective: 2.5

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Answer Key
Testname: UNTITLED2

85) FALSE
ID: TestBank 2.3-9
Diff: 0
Objective: 2.5

86) TRUE
ID: TestBank 2.3-10
Diff: 0
Objective: 2.5

87) FALSE
ID: TestBank 2.3-11
Diff: 0
Objective: 2.6

88) TRUE
ID: TestBank 2.3-12
Diff: 0
Objective: 2.6

89) FALSE
ID: TestBank 2.3-13
Diff: 0
Objective: 2.7

90) FALSE
ID: TestBank 2.3-14
Diff: 0
Objective: 2.7

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91) 505
ID: TestBank 2.4-1
Diff: 0
Objective: 2.2

92) 8.32
ID: TestBank 2.4-2
Diff: 0
Objective: 2.2

93) 25 200
ID: TestBank 2.4-3
Diff: 0
Objective: 2.2

94) 58.5
ID: TestBank 2.4-4
Diff: 0
Objective: 2.2

95) 0.00341
ID: TestBank 2.4-5
Diff: 0
Objective: 2.2

96) 3
ID: TestBank 2.4-6
Diff: 0
Objective: 2.2

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97) 2
ID: TestBank 2.4-7
Diff: 0
Objective: 2.2

98) 5
ID: TestBank 2.4-8
Diff: 0
Objective: 2.2

99) 1
ID: TestBank 2.4-9
Diff: 0
Objective: 2.2

100) 4
ID: TestBank 2.4-10
Diff: 0
Objective: 2.2

101) 4
ID: TestBank 2.4-11
Diff: 0
Objective: 2.2

102) 2
ID: TestBank 2.4-12
Diff: 0
Objective: 2.2

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103) 3
ID: TestBank 2.4-13
Diff: 0
Objective: 2.2

104) cubic meter
ID: TestBank 2.4-14
Diff: 0
Objective: 2.1

105) gram
ID: TestBank 2.4-15
Diff: 0
Objective: 2.1

106) 29.2 g
ID: TestBank 2.4-16
Diff: 0
Objective: 2.5

107) mass, volume
ID: TestBank 2.4-17
Diff: 0
Objective: 2.7

108) 57 mL
ID: TestBank 2.4-18
Diff: 0
Objective: 2.7

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109) 290. g of gold
ID: TestBank 2.4-19
Diff: 0
Objective: 2.7

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