# *Chemistry for Engineering Students 4th edition Brown Test Bank*

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**Chapter 02**

1. All polymers form rigid solids at standard temperature and pressure.
   a. True
   b. False

*ANSWER:* False
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* Conducting Polymers
*DATE CREATED:* 12/23/2013 1:59 PM
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2. The nucleus of an atom is comprised of protons and electrons.
   a. True
   b. False

*ANSWER:* False
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* Atomic Structure and Mass
*DATE CREATED:* 12/23/2013 1:59 PM
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3. Cations are particles with fewer electrons than protons.
   a. True
   b. False

*ANSWER:* True
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* Ions
*DATE CREATED:* 12/23/2013 1:59 PM
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4. A horizontal row of elements on the periodic table is referred to as a family.
   a. True
   b. False

*ANSWER:* False
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False

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*TOPICS:* The Periodic Table
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5. The systematic name of $V_2O_5$ is vanadium pentoxide.
   a. True
   b. False
*ANSWER:* True
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
*DATE CREATED:* 12/23/2013 1:59 PM
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6. The systematic name of $AlCl_3$ is aluminum chloride.
   a. True
   b. False
*ANSWER:* True
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
*DATE CREATED:* 12/23/2013 1:59 PM
*DATE MODIFIED:* 1/19/2018 12:55 AM

7. The chemical formula for nitrite is $NO_3^-$.
   a. True
   b. False
*ANSWER:* False
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
*DATE CREATED:* 12/23/2013 1:59 PM
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8. Elements in a periodic family tend to combine with the same number of hydrogen atoms.
   a. True
   b. False
*ANSWER:* True

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*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* The Periodic Table
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9. Free radicals, such as Cl·, tend to be very reactive.
   a. True
   b. False
*ANSWER:* True
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* Polyethylene
*DATE CREATED:* 12/23/2013 1:59 PM
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10. Silicon tetrachloride plays an important role in the production of semiconductors.
   a. True
   b. False
*ANSWER:* True
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* Inorganic and Organic Chemistry
*DATE CREATED:* 12/23/2013 1:59 PM
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11. $CH_3{-}O{-}CH_3$ is an example of an alcohol.
   a. True
   b. False
*ANSWER:* False
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* Inorganic and Organic Chemistry
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12. Chlorine forms four oxyanions.
    a. True
    b. False

*ANSWER:* True
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
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13. The average mass of an atom is determined by:
    a. adding the number of protons and electrons and dividing it by the number of neutrons.
    b. averaging the masses of each isotope.
    c. calculating the weighted average of all stable isotopic masses.
    d. adding the number of protons and neutrons and dividing it by the number of electrons.

*ANSWER:* c
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Atomic Structure and Mass
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14. Ions of opposite charges attract one another. This attraction is governed by _____.
    a. Graham's law
    b. Henry's law
    c. Kirchoff's law
    d. Coulomb's law

*ANSWER:* d
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Ions
*DATE CREATED:* 12/23/2013 1:59 PM
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15. In the modern periodic table, the densest element is found in _____.
    a. period 2
    b. period 4
    c. period 6

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d. period 8

*ANSWER:* c
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* The Periodic Table
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16. Group II elements (Be, Mg, Ca, Sr, and Ba) are commonly referred to as _____.

a. alkali metals
b. alkaline earth metals
c. halogen metals
d. lanthanide metals

*ANSWER:* b
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* The Periodic Table
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17. The vertical columns in the periodic table are called _____.

a. groups
b. periods
c. classes
d. sections

*ANSWER:* a
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* The Periodic Table
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18. In the context of the modern periodic table, _____ are examples of transition metals.

a. Fe and Zn
b. Sb and I
c. Pm and Gd
d. Al and Ga

*ANSWER:* a

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*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* The Periodic Table
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19. _____ are examples of actinides.
    a. Ti and Cr
    b. U and Np
    c. Sm and Er
    d. Kr and Xe

*ANSWER:* b
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* The Periodic Table
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20. Which of the following pairs of elements can be classified as noble gases?
    a. H and He
    b. Na and K
    c. Kr and Ar
    d. N and O

*ANSWER:* c
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* The Periodic Table
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21. Uranium, plutonium, and neptunium are classified as _____.
    a. alkali metals
    b. alkaline earth metals
    c. lanthanides
    d. actinides

*ANSWER:* d
*POINTS:* 1
*DIFFICULTY:* Easy

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*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* The Periodic Table
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22. Alkali metal cations carry a charge of _____.
    a. 1+
    b. 2+
    c. 2−
    d. 1−

*ANSWER:* a
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* The Periodic Table
*DATE CREATED:* 12/23/2013 1:59 PM
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23. $CaCl_2$ is an example of a(n) _____.
    a. covalent compound
    b. formula unit
    c. molecular compound
    d. organic acid

*ANSWER:* b
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Compounds and Chemical Bonds
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24. The molecular formula for lead nitrate is _____.
    a. $PbN_3$
    b. $Pb(NO_2)_3$
    c. $Pb(NO_3)_2$
    d. $PNO_3$

*ANSWER:* c
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice

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*TOPICS:* Chemical Nomenclature
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25. The molecular formula for ammonium chloride is _____.

a. $NH_3Cl$
b. $NH_2Cl_2$
c. $NH_4Cl$
d. $NH_2Cl_3$

*ANSWER:* c
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
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26. The molecular formula for strontium chloride is _____.

a. $SrCl_2$
b. $StCl_2$
c. $SrClO_2$
d. SrCl

*ANSWER:* a
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
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27. The molecular formula for iron(II) bromide is _____.

a. $FeBr_3$
b. $FeBr_2$
c. $I_2Br_2$
d. $IBr_3$

*ANSWER:* b
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice

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*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
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28. The molecular formula for dinitrogen pentoxide is _____.

a. $N_2O_5$
b. $(NO_5)_2$
c. $2NO_5$
d. $N_2P_5O$

*ANSWER:* a
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
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29. The molecular formula for rubidium chloride is _____.

a. $RuCl_2$
b. $RbCl_2$
c. RuCl
d. RbCl

*ANSWER:* d
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
*DATE CREATED:* 12/23/2013 1:59 PM
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30. The molecular formula for calcium phosphate is _____.

a. $Ca_3P_2$
b. $CaPO_3$
c. $Ca_3(PO_4)_2$
d. $Ca_2(PO_4)_3$

*ANSWER:* c
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice

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*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
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31. The molecular formula for tin(IV) oxide is _____.

a. $TiO_2$
b. $SnO_2$
c. $SnO_4$
d. $TiO_4$

*ANSWER:* b
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
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32. The systematic (IUPAC) name of $MgI_2$ is _____.

a. manganese iodide
b. manganese diiodide
c. magnesium iodide
d. magnesium (II) iodide

*ANSWER:* c
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
*DATE CREATED:* 12/23/2013 1:59 PM
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33. The systematic (IUPAC) name of $Sr(NO_2)_2$ is _____.

a. strontium nitrate
b. strontium dinitrate
c. sirium nitrate
d. strontium nitrite

*ANSWER:* d
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False

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*TOPICS:* Chemical Nomenclature
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34. The systematic (IUPAC) name of CuO is _____.
    a. copper (I) oxide
    b. copper (II) oxide
    c. copper oxide
    d. copper (II) hydroxide

*ANSWER:* b
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
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35. The systematic (IUPAC) name of $NaClO_4$ is _____.
    a. sodium perchlorate
    b. sodium chlorate
    c. sodium hypochlorate
    d. sodium chloride tetraoxide

*ANSWER:* a
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Chemical Nomenclature
*DATE CREATED:* 12/23/2013 1:59 PM
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36. How many neutrons are present in $^{201}Hg$?
    a. 201
    b. 80
    c. 283
    d. 121

*ANSWER:* d
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Atomic Structure and Mass
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37. How many electrons are present in $^{40}Ca^{2+}$?

a. 40
b. 38
c. 20
d. 18

*ANSWER:* d
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Ions
*DATE CREATED:* 12/23/2013 1:59 PM
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38. Which of the following chemical species contains 9 protons?

a. $^{16}O$
b. $^{19}F$
c. $^{35}S$
d. $^{32}P$

*ANSWER:* b
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Atomic Structure and Mass
*DATE CREATED:* 12/23/2013 1:59 PM
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39. If an ion contains 33 protons, 39 neutrons, and 34 electrons, the ion is _____.

a. $^{73}Se^{1-}$
b. $^{72}As^{1-}$
c. $^{67}Y^{1+}$
d. $^{73}Se^{1+}$

*ANSWER:* b
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Ions
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40. Consider $^{235}U$. How many neutrons are present in this nuclide?

a. 143
b. 235
c. 92
d. 177

*ANSWER:* a
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Atomic Structure and Mass
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41. When electrons are shared between pairs of atoms rather than donated from one atom to another or mobile across an entire lattice, we have _____ bonding.

*ANSWER:* covalent
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Completion
*HAS VARIABLES:* False
*TOPICS:* Compounds and Chemical Bonds
*DATE CREATED:* 12/23/2013 1:59 PM
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42. The atomic mass of an atom is the number of _____ in that particular atom.

*ANSWER:* protons and neutrons
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Completion
*HAS VARIABLES:* False
*TOPICS:* Atomic Structure and Mass
*DATE CREATED:* 12/23/2013 1:59 PM
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43. An _____ provides the relative ratio between the numbers of atoms of the different elements present in a compound.

*ANSWER:* empirical formula
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Completion
*HAS VARIABLES:* False
*TOPICS:* Compounds and Chemical Bonds

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44. The ratio of isotopes for a given element can be measured instrumentally using a _____.
*ANSWER:* mass spectrometer
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Completion
*HAS VARIABLES:* False
*TOPICS:* Atomic Structure and Mass
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45. Atoms of the same element that have different numbers of neutrons are called _____.
*ANSWER:* isotopes
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Completion
*HAS VARIABLES:* False
*TOPICS:* Atomic Structure and Mass
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46. Potassium has two stable isotopes: $^{39}K$, which has a mass of 38.96 amu and makes up 93.26 % of the natural potassium found; and $^{41}K$, with a mass of 40.96 amu. What is the atomic weight of potassium?
*ANSWER:* 39.09 amu
*POINTS:* 1
*DIFFICULTY:* Hard
*QUESTION TYPE:* Subjective Short Answer
*HAS VARIABLES:* False
*TOPICS:* Atomic Structure and Mass
*DATE CREATED:* 1/19/2018 4:09 AM
*DATE MODIFIED:* 1/19/2018 4:36 AM

47. Anions are negatively charged species that contain fewer electrons than protons.
a. True
b. False
*ANSWER:* False
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* Ions
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48. Barium is an alkali metal.
    a. True
    b. False

*ANSWER:* False
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* True / False
*HAS VARIABLES:* False
*TOPICS:* The Periodic Table
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49. Which of the following is a metalloid?
    a. Boron
    b. Lithium
    c. Uranium
    d. Tungsten

*ANSWER:* a
*POINTS:* 1
*DIFFICULTY:* Easy
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* The Periodic Table
*DATE CREATED:* 1/19/2018 4:14 AM
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50. Which of the following isotopes has the highest number of neutrons?
    a. $^{85}Rb$
    b. $^{28}Si$
    c. $^{79}Br$
    d. $^{83}Kr$

*ANSWER:* a
*POINTS:* 1
*DIFFICULTY:* Moderate
*QUESTION TYPE:* Multiple Choice
*HAS VARIABLES:* False
*TOPICS:* Atomic Structure and Mass
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